

# The Crystal Structure of Tetrapotassium Dihydrogendecaoxodiiodate(VII) Octahydrate

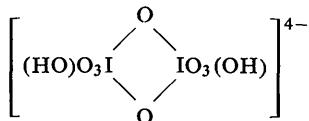
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The crystal structure of tetrapotassium dihydrogendecaoxodiiodate (VII) octahydrate,  $K_4H_2I_2O_{10} \cdot 8H_2O$ , has been determined and refined.

The binuclear anion



can be recognized in the structure. It is formed by two octahedral groups with an edge in common, i.e. with two oxygen atoms shared between two iodine atoms. Distances within the binuclear anion are  $I-O_{av}^* = 1.996 \pm 0.021 \text{ \AA}$  for shared oxygen atoms, including that of the probable hydroxyl group, and  $I-O_{av} = 1.807 \pm 0.011 \text{ \AA}$  for unshared oxygen atoms.

The anions are interposed between  $K^+$  cations and water molecules. The water molecules form a shell around the anion to which they are bound by hydrogen bonds. Each water molecule is bound to two different anions, and a three-dimensional network of hydrogen bonds branches all over the crystal.

The potassium ions are surrounded each by eight water molecules or oxygen atoms but their environments are not equivalent.

## Introduction

The molecular and crystal structures of tetraoxoiodates(VII) (i.e. metaperiodates) (Hazlewood, 1938; Jaeger & Beintema, 1933) and of some orthoperiodates (Helmholz, 1937; Ferrari, Curti & Cavalca, 1944) are already known. The tetrahedral group  $[IO_4]^-$  is present in the former compounds, the octahedral group  $[IO_6]^{5-}$  in the latter. A distorted octahedral group has been found in orthoperiodic acid,  $(HO)_5IO$  (Feikema, 1961, 1963). The existence of other oxoiodic(VII) acids or oxoiodates(VII) has been demonstrated or postulated by several authors (Rammelsberg, 1868; Blomstrand, 1886; Kimmins, 1889; Rae, 1931; Partington & Bahl, 1934; Pačesová & Hauptmann, 1963). In particular Siebert (1960a, b), from an examination of the infrared spectra of several oxoiodates(VII) infers the existence of the so-called mesoperiodic anion  $[IO_5]^{3-}$ , which should be present, for instance, in the salt  $K_2HIO_5 \cdot 4H_2O$ . The formula of this salt can also be written  $K_4I_2O_9 \cdot 9H_2O$  (Hill, 1928) or  $K_4H_2I_2O_{10} \cdot 8H_2O$ . The crystal structure of this compound has now been investigated and the last formulation has been proved to be correct.

## Experimental

Crystals of the salt  $K_4H_2I_2O_{10} \cdot 8H_2O$  can be obtained, either by evaporating or by cooling, from solutions of potassium periodate in concentrated aqueous potassium hydroxide. Its crystallographic properties were

studied by Groth (1908), who assigned them to the triclinic system, pinacoidal class.

A crystal elongated parallel to [001], according to the orientation of Groth, has been used to determine the unit-cell constants by the roentgenographic method. The crystal has been rotated around the elongation axis, which is parallel to  $c$ . The linear constants have been measured on the photographs and the angular constants have been calculated from values given by Groth (1908) after transformation by the matrix 101/010/001. The resulting unit-cell constants are:

$$\begin{array}{ll} a = 7.06 \pm 0.01 \text{ \AA} & \alpha = 97^\circ 47' \\ b = 10.57 \pm 0.01 & \beta = 121^\circ 30' \\ c = 7.56 \pm 0.01 & \gamma = 81^\circ 55' \end{array}$$

One stoichiometric unit  $K_4H_2I_2O_{10} \cdot 8H_2O$  is contained in the unit cell. ( $\rho_{\text{calc}} = 2.500 \text{ g.cm}^{-3}$ ,  $\rho_{\text{obs}} = 2.509 \text{ g.cm}^{-3}$ ). The cell reported above can be transformed to the conventional cell (*International Tables for X-ray Crystallography*, 1959) by the matrix 101/010/ $\bar{1}00$ .

The space group  $P\bar{1}$  was assumed, in accordance with the morphological properties.

For the structural determinations, a crystal rotating around [001] was reduced to a cylindrical shape by fining it by a paper strip wetted by alkali solution. The intensities of the integrated reflexions  $hk0$ ,  $hk1$ ,  $hk2$ ,  $hk3$ ,  $hk4$ ,  $hk5$  have been recorded on multiple films in a Weissenberg camera and then measured by a microphotometer. Cu  $K\alpha$  radiation was used.

Lorentz and polarization factors were calculated. The transmission factors were obtained by interpolation from data given by Bond (1959) for cylindrical specimens ( $\mu R = 1.9$ ).

The atomic form factors were calculated by the Forsyth & Wells (1959) function, using the improved constants by Moore (1963). All the corrections, structure factors, three-dimensional Fourier functions, differential syntheses and anisotropic temperature factor refinement were calculated with an Olivetti Elea 6001/S computer.

### Determination of structure

The first step of the determination of the structure was the interpretation of the Patterson projection  $P(uv)$ . Vectors I-I, K(1)-I and K(2)-I could be identified and satisfactory agreement between observed and calculated structure factors  $F(hk0)$  was obtained. The agreement was improved by calculating a few electron density projections  $\rho_0(xy)$ . The  $z$  coordinates of I, K(1) and K(2) were found approximately by generalized Patterson functions  $cP_1(uv)$  and  $sP_1(uv)$ . The remaining atoms were identified in the three-dimensional  $\rho_0(xyz)$  (Fig. 1).

The refinement started at  $R=0.164$ . The coordinates were refined by the differential synthesis method (Booth, 1946); for the refinement of the anisotropic

thermal parameters the method described by Nardelli, Fava & Giraldi (1963) has been followed. The two refinements were applied alternately. The final atomic positions with their e.s.d.'s are reported in Table 1. (Final  $R_{hkl}=0.099$ , observed reflexions only). The anisotropic thermal parameters are quoted in Table 2. The

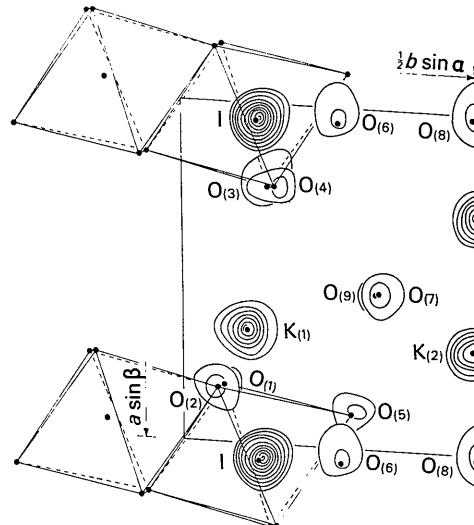


Fig. 1.  $\rho(xyz)$ , viewed down [001]. Composite map, arbitrary intervals. Octahedra forming binuclear anions are also drawn.

Table 1. Final atomic coordinates

Atom	$x/a$	$\sigma(x/a)$	$y/b$	$\sigma(y/b)$	$z/c$	$\sigma(z/c)$
I	0.0503	0.0001	0.1261	0.0001	0.4512	0.0002
K(1)	0.6699	0.0006	0.1117	0.0004	0.7654	0.0007
K(2)	0.6957	0.0006	0.4931	0.0004	0.8454	0.0007
O(1)	0.8300	0.0023	0.0721	0.0014	0.1601	0.0022
O(2)	0.8424	0.0014	0.0611	0.0006	0.5147	0.0018
O(3)	0.2436	0.0022	0.1445	0.0011	0.3722	0.0022
O(4)	0.2373	0.0017	0.1558	0.0010	0.7238	0.0023
O(5)	0.8976	0.0021	0.2833	0.0008	0.4025	0.0025
O(6)	0.0433	0.0023	0.2637	0.0014	0.9733	0.0029
O(7)	0.5404	0.0021	0.3341	0.0011	0.9953	0.0027
O(8)	0.0120	0.0023	0.4967	0.0011	0.7319	0.0028
O(9)	0.5414	0.0022	0.3302	0.0012	0.4844	0.0033

Table 2. Anisotropic thermal parameters\* ( $\text{\AA}^2$ )

	$B_{11}$	$B_{22}$	$B_{33}$	$B_{12}$	$B_{13}$	$B_{23}$
I	1.302	1.006	1.418	-0.112	0.488	0.137
K(1)	1.848	2.195	1.827	0.092	0.552	0.195
K(2)	2.457	1.940	2.714	-0.138	1.024	0.195
O(1)	3.092	3.236	2.035	-0.923	0.210	0.840
O(2)	1.496	1.387	2.343	0.010	0.236	0.434
O(3)	2.191	1.400	2.480	-0.759	0.703	-0.429
O(4)	2.228	1.727	3.255	0.120	0.946	0.730
O(5)	3.254	1.231	3.277	-0.931	1.484	0.186
O(6)	2.355	2.505	3.469	0.146	0.867	0.148
O(7)	2.825	1.798	3.697	-0.269	1.326	0.521
O(8)	4.977	2.582	4.869	-0.410	2.428	-0.219
O(9)	2.093	2.718	3.442	0.587	0.543	0.898

\* In the last cycle, for each atomic species I, K and O, the average shift,  $|\Delta B_{ij}|$ , and the maximum shift,  $|\Delta B_{ij}|_{\max}$ , were:

for I:  $|\Delta B_{ij}|=0.019$  and  $|\Delta B_{ij}|_{\max}=0.062$

for K:  $|\Delta B_{ij}|=0.020$  and  $|\Delta B_{ij}|_{\max}=0.060$

for O:  $|\Delta B_{ij}|=0.031$  and  $|\Delta B_{ij}|_{\max}=0.179$

The averages have been made over all the parameter shifts of all the atoms of the same species.

Table 3. Observed and calculated structure factors

$\frac{h}{l}$	$\frac{k}{l}$	$\frac{l}{l}$	$10F_0$	$10F_C$	$\frac{h}{l}$	$\frac{k}{l}$	$\frac{l}{l}$	$10F_0$	$10F_C$	$\frac{h}{l}$	$\frac{k}{l}$	$\frac{l}{l}$	$10F_0$	$10F_C$	$\frac{h}{l}$	$\frac{k}{l}$	$\frac{l}{l}$	$10F_0$	$10F_C$	$\frac{h}{l}$	$\frac{k}{l}$	$\frac{l}{l}$	$10F_0$	$10F_C$					
1	0	0	819	962	2	7	0	—	32	4	1	1	671	560	8	4	1	—	238	6	8	1	—	87	7	0	2	230	272
2	0	0	471	434	3	7	0	588	591	4	1	1	356	403	6	8	1	—	93	8	0	2	211	291	4	4	2	556	573
3	0	0	631	609	3	7	0	169	155	5	1	1	293	261	6	8	1	—	290	6	9	1	—	268	4	4	2	318	274
4	0	0	283	274	4	7	0	318	331	5	1	1	241	252	0	5	1	358	461	0	1	2	777	946	4	4	2	435	384
5	0	0	194	162	4	7	0	—	76	5	1	1	340	277	0	5	1	400	348	0	1	2	226	222	5	4	2	260	242
6	0	0	—	5	5	7	0	372	394	5	1	1	491	525	1	5	1	589	586	0	9	1	334	348	1	1	2	1560	1281
7	0	0	181	169	5	7	0	301	302	6	1	1	320	263	1	5	1	485	615	0	9	1	325	272	1	1	2	723	863
8	0	0	181	169	6	7	0	—	10	6	1	1	320	263	1	5	1	320	263	0	9	1	325	272	1	1	2	723	863
9	1	0	635	758	6	7	0	278	314	6	1	1	451	367	1	5	1	319	257	1	9	1	219	282	1	1	2	789	906
10	1	0	587	711	6	7	0	—	10	6	1	1	377	404	2	5	1	—	10	1	9	1	345	432	1	1	2	233	277
11	1	0	684	787	0	8	0	746	775	7	1	1	200	193	2	5	1	1032	1223	1	9	1	487	414	2	1	2	574	572
12	1	0	440	376	1	8	0	327	295	7	1	1	120	135	2	5	1	418	405	1	9	1	87	95	2	1	2	673	793
13	1	0	603	649	1	8	0	657	718	7	1	1	344	331	3	5	1	184	167	2	9	1	171	196	2	1	2	451	454
14	1	0	—	81	2	8	0	461	446	8	1	1	77	75	3	5	1	638	594	2	9	1	390	336	3	1	2	142	186
15	1	0	314	305	2	8	0	232	191	8	1	1	—	229	3	5	1	388	353	2	9	1	166	146	3	1	2	641	686
16	1	0	465	446	3	8	0	646	667	8	1	1	373	363	3	5	1	121	168	3	9	1	419	369	6	4	2	314	374
17	1	0	421	417	3	8	0	470	510	8	1	1	209	189	4	5	1	362	379	3	9	1	482	451	—	—	—	—	55
18	1	0	489	405	4	8	0	165	151	0	2	1	170	106	4	5	1	209	189	3	9	1	558	470	4	1	2	27	0
19	1	0	605	570	4	8	0	191	205	0	2	1	213	152	4	5	1	654	590	3	9	1	155	142	4	1	2	509	518
20	1	0	487	439	5	8	0	164	160	1	2	1	461	415	4	5	1	658	740	4	9	1	—	32	4	1	2	701	698
21	1	0	351	337	5	8	0	—	40	1	2	1	589	660	5	5	1	271	270	4	9	1	254	291	4	1	2	887	942
22	1	0	332	302	6	8	0	—	74	1	2	1	464	358	5	5	1	412	357	5	5	1	456	397	5	5	1	—	61
23	1	0	108	96	6	7	0	—	—	1	2	1	857	1005	5	5	1	350	319	4	9	1	315	312	5	5	1	344	370
24	2	0	269	196	1	9	0	268	213	2	2	1	874	1146	6	5	1	366	366	5	9	1	192	213	5	5	1	344	370
25	2	0	224	236	1	9	0	279	289	2	2	1	189	159	6	5	1	415	377	5	9	1	191	193	2	1	2	280	305
26	2	0	279	198	2	9	0	—	40	2	2	1	394	346	6	5	1	142	103	5	9	1	405	350	2	1	2	301	282
27	2	0	744	874	2	9	0	543	608	3	2	1	401	455	7	5	1	347	369	6	9	1	329	300	6	1	2	43	44
28	2	0	193	177	3	9	0	127	86	3	2	1	512	429	7	5	1	406	451	0	10	1	—	9	7	1	2	168	94
29	2	0	361	344	3	9	0	416	422	3	2	1	408	432	7	5	1	408	432	0	10	1	213	201	7	1	2	303	349
30	2	0	720	702	4	9	0	253	193	4	2	1	373	268	6	2	1	361	367	1	10	1	126	155	8	1	2	239	239
31	2	0	436	450	4	9	0	286	258	4	2	1	507	540	0	6	1	—	34	5	6	1	—	24	1	10	1	192	142
32	2	0	357	359	5	9	0	278	300	4	2	1	675	546	0	6	1	378	346	6	5	1	204	192	2	1	2	302	372
33	2	0	231	311	5	9	0	199	253	4	2	1	787	919	1	6	1	203	146	2	10	1	512	431	3	10	1	714	838
34	2	0	231	311	5	9	0	199	253	5	2	1	395	337	1	6	1	447	523	2	10	1	178	212	0	2	2	199	200
35	2	0	721	685	6	9	0	217	281	5	2	1	677	742	1	6	1	181	175	2	10	1	497	564	2	10	1	223	256
36	2	0	264	230	6	9	0	—	—	5	2	1	582	463	1	6	1	619	593	2	10	1	176	132	3	5	2	142	120
37	2	0	506	474	6	9	0	93	0	5	10	0	349	397	8	2	1	477	520	5	2	1	395	405	4	10	1	654	576
38	2	0	307	230	1	10	0	248	237	8	2	1	537	503	2	6	1	425	478	3	10	1	337	357	3	2	2	228	175
39	2	0	210	180	1	10	0	—	86	6	2	1	361	367	2	6	1	85	52	3	10	1	350	402	3	2	2	1384	1442
40	3	0	921	1034	2	10	0	413	430	6	2	1	466	380	2	6	1	642	656	3	10	1	229	182	4	2	2	490	504
41	3	0	517	528	3	10	0	129	145	6	2	1	267	273	3	6	1	303	346	3	10	1	486	435	3	10	1	797	756
42	3	0	313	800	3	10	0	422	423	7	2	1	129	138	3	6	1	417	399	3	10	1	486	435	3	10	1	201	257
43	3	0	327	330	4	10	0	366	406	7	2	1	250	250	3	6	1	366	351	4	10	1	280	295	3	5	2	235	136
44	3	0	460	435	4	10	0	326	356	7	2	1	272	279	3	6	1	395	351	4	10	1	282	219	3	5	2	229	260
45	3	0	1101	1110	5	10	0	349	376	8	2	1	109	109	4	6	1	381	482	5	11	1	317	346	3	6	2	288	557
46	3	0	423	347	5	10	0	—	71	5	2	1	397	219	0	7	1	460	511	2	11	1	571	482	4	6	2	228	175
47	3	0	823	813	5	10	0	253	281	5	3	1	166	136	0	7	1	474	524	2	11	1	571	482	5	6	2	192	265
48	4	0	298	298	0	12	0	253	281	5	3	1	332	339	0	7	1	301	206	3	10	1	321	560	4	6	2	265	268
49	4	0	537	509	1	12	0	344	374	3	3	1	364	407	0	7	1	321	321	3	11	1	301	206	5	6	2	226	223
50	4	0	342	312	2	12	0	283	312	4	3	1	824	766	1	7	1	151	101	3	11	1	437	388	3	11	1	286	323
51	4	0	—	30	3	12	0	188	186	4	3	1	126	108	1	7	1	786	730	4	11	1	364	441	4	11	1	267	307
52	4	0	42	42	3	12	0	173	173	4	3	1	455	492	2	7	1	164	211	4	11	1	254	292	4	11			

Table 3 (cont.)

$\hbar$	$k$	$l$	$10P_0$	$10P_2$	$\hbar$	$k$	$l$	$10P_0$	$10P_2$	$\hbar$	$k$	$l$	$10P_0$	$10P_2$	$\hbar$	$k$	$l$	$10P_0$	$10P_2$	$\hbar$	$k$	$l$	$10P_0$	$10P_2$								
2	8-	2	584	566	4-	0	3	251	231-	4	4	3	378	322	3-	0	3	91-	2	1	4	275	255	1	5	4	523	562-				
2-	-	2	98-	54-	4-	0	3	425	316-	4-	4	3	304	215-	4-	0	3	255	263-	2	1	4	286	320	1-	5	4	170	180-			
3	8-	2	283	245	8-	0	3	320	265-	4-	4	3	176	157-	4-	0	3	99	89	2	1	4	117	25	1-	5	4	170	169-			
3-	-	2	51-	50-	0	3	811	811	4-	4	3	138	322-	4-	0	3	163	216	2-	1	4	1204-	—	1-	5	4	434	422-				
3-	8-	2	491	479	6-	0	3	343	301	5-	4	3	475	462	4-	0	3	320	322	3	1	4	472	478	2-	5	4	335	324-			
4-	8-	2	236	176-	7-	0	3	217	217	5-	4	3	464	420-	5-	0	3	430	450	3-	1	4	191	204	2-	5	4	568	523-			
4-	8-	2	331	292	8-	0	3	431	438	5-	4	3	263	314	5-	0	3	464	476	3-	1	4	407	403	2-	5	4	167	118-			
4-	8-	2	88-	234	196	-	-	-	-	5-	4	3	472	467-	6-	0	3	328	350	3-	1	4	396	450-	2-	5	4	333	334-			
4-	8-	2	24-	0	1	3	858	1004-	6-	4	3	434	432-	7-	0	3	321	292	4-	1	4	—	66-	3-	5	4	378	329-				
5-	8-	2	187	214	0-	1	3	389	428	7-	4	3	218	233-	4-	1	4	176	170	3-	5	4	180	157-	1-	10	4	341	383-			
5-	8-	2	95-	1	1	3	701	814-	7-	4	3	372	364-	4-	1	4	176	170	3-	5	4	167	164-	1-	10	4	323	329-				
6-	8-	2	388	364-	1-	1	3	592	714-	8-	4	3	231	231-	0	9	3	379	383-	5-	1	4	346	360	1-	10	4	278	322-			
6-	8-	2	331	402-	1-	1	3	81	48-	4-	0	3	350	347-	1-	0	3	351	353	5-	1	4	432	336-	4-	5	4	206	209			
7-	8-	2	416	403-	1-	1	3	250	192-	1-	0	3	657	658-	1-	0	3	464	483-	5-	1	4	753	693-	4-	5	4	580	686			
7-	2	272	250-	2	1	3	657	658-	2-	1	3	283	291-	0	5	3	843	932	78-	6-	1	279	219-	5-	4	143	154	2-	10	4	457	433-
7-	2	1-	3	562	605-	0	5	3	107	65-	1-	0	3	—	26	6-	1	4	320	283-	5-	4	256	284	3-	10	4	343	329-			
0	9-	2	587	521	2-	1	3	240	339	1-	5	3	316	309	2-	0	3	438	403-	7-	1	4	413	366-	6-	5	4	127	117			
0	9-	2	40-	3	1	3	254	235-	1-	5	3	537	595	2-	0	3	273	271-	7-	1	4	59-	59-	6-	5	4	222	207				
1-	9-	2	456	403	3-	1	3	278	245-	1-	5	3	76	44-	2-	0	3	374	384-	8-	1	4	267	272-	7-	5	4	366	392			
1-	9-	2	434	375	3-	1	3	210	202-	1-	5	3	268	268-	2-	0	3	305	305	8-	1	4	162	135-	7-	5	4	230	229-			
1-	9-	2	273	213	3-	1	3	654	621	2-	0	3	242	232-	9-	1	4	203	194-	8-	5	4	45	95-	8-	5	4	93	102			
1-	9-	2	288	244	4-	1	3	124	119-	2-	0	3	293	293-	3-	0	3	293	308-	2-	2	4	770	786	2-	6	4	495	484-			
2-	9-	2	524	438	4-	1	3	232	208-	3-	0	3	662	731-	3-	0	3	522	529	0	2	4	431	434	0	6	4	485	498-			
2-	9-	2	370	367	4-	1	3	566	545	3-	0	3	352	316	4-	0	3	114	126-	0	2	4	554	636-	0	6	4	415	431			
2-	9-	2	75-	5	1	3	138	143-	4-	0	3	540	526	4-	0	3	317	335	1-	2	4	737	782	1-	6	4	423	369-				
3-	9-	2	70	51-	1	3	198	149-	3-	0	3	364	352	5-	0	3	37	37	1-	2	4	507	535	1-	6	4	532	569-				
3-	9-	2	358	312	5-	1	3	506	504-	3-	0	3	486	564-	3-	0	3	363	353	1-	2	4	461	445-	1-	6	4	524	547-			
3-	9-	2	281	239	5-	1	3	479	423	4-	0	3	170	117	6-	0	3	364	345	1-	2	4	533	526-	2-	11	4	476	250-			
3-	9-	2	511	465-	6-	1	3	186	191-	4-	0	3	255	277-	7-	0	3	293	203	2-	2	4	146	166-	2-	6	4	559	318-			
4-	9-	2	199	176-	6-	1	3	383	358	4-	0	3	207	226-	7-	0	3	433	433-	2-	2	4	770	786	2-	6	4	495	484-			
4-	9-	2	275	277	7-	1	3	253	260	6-	0	3	514	569-	7-	0	3	295	292	2-	2	4	247	234-	2-	6	4	301	298-			
4-	9-	2	445	458	7-	1	3	113	138-	5-	0	3	305	261-	8-	0	3	295	292	3-	2	4	382	522-	2-	6	4	909	940			
4-	9-	2	515	447-	8-	1	3	282	273	5-	0	3	305	261-	9-	0	3	303	303	3-	2	4	411	485-	3-	6	4	447	470-			
5-	9-	2	78	88-	8-	1	3	138	129-	5-	0	3	231	260	10-	0	3	277	260	3-	2	4	473	566-	3-	6	4	322	338-			
5-	9-	2	401	356-	6-	0	3	360	334-	5-	0	3	607	651-	10-	0	3	403	383-	3-	2	4	144	103-	3-	6	4	122	106-			
6-	9-	2	227	254-	0	0	3	1049	1229-	6-	0	3	313	295-	10-	0	3	431	393	4-	0	4	138	110-	0	0	4	21	0			
0	10-	2	61	61	0	2	3	459	463	7-	5	3	95	99-	10-	0	3	203	191-	19-	0	4	424	384-	0	12	4	86	96-			
0	10-	2	464	421-	1-	2	3	325	348-	8-	5	3	237	227-	20-	0	3	223	226-	4-	2	4	143	115-	0	6	4	284	311-			
1-	10-	2	550	489-	6-	2	3	351	355-	5-	0	3	630	604	10-	0	3	286	262-	4-	2	4	397	375-	4-	6	4	303	274-			
1-	10-	2	572	182-	6-	2	3	349	349-	6-	0	3	324	314-	10-	0	3	303	299-	5-	2	4	143	125-	2-	12	4	110	123-			
1-	10-	2	257	264-	6-	2	3	125	137-	1-	0	3	217	212-	10-	0	3	386	381-	5-	2	4	329	273-	2-	12	4	222	223-			
2-	10-	2	365	318	6-	2	3	201	178-	6-	0	3	329	349-	10-	0	3	253	229-	7-	2	4	181	144-	2-	12	4	87	94-			
2-	10-	2	478-	478-	7-	2	3	86	86-	6-	0	3	328	326-	10-	0	3	277	268-	7-	2	4	194	154-	2-	13	4	345	345-			
2-	10-	2	519	478-	7-	2	3	86	86-	6-	0	3	328	326-	10-	0	3	277	268-	7-	2	4	195-	195-	2-	13	4	193	193-			
3-	10-	2	156	175-	7-	2	3	153	134-	6-	0	3	312	305-	10-	0	3	277	268-	7-	2	4	181	144-	2-	13	4	345	345-			
3-	10-	2	217	193-	7-	2	3	153	134-	6-	0	3	312	305-	10-	0	3	269	291-	7-	2	4	193	184-	2-	13	4	99	98-			
3-	10-	2	561	497-	4-	2	3	166	175-	3-	0	3	440	427-	11-	0	3	478	503-	0	3	4	221	236-	1-	7	4	225	221-			
4-	10-	2	473	461-	5-	2	3	245	187-	6-	0	3	374	368-	11-	0	3	399	357-	1-	3	4	316	270-	2-	7	4	233	283-			
4-	10-	2	381	344-	5-	2	3	209	207-	6-	0	3	381	383-	11-	0	3	357	381-	1-	3	4	406	524-	2-	7	4	266	270-			
5-	10-	2	192	192-	5-	2	3	309	907-	6-	0	3	330	345-	11-	0	3	315	315-	2-	7	4	267	260-	2-	7	4	266	270-			
5-	10-	2	233	223-	6-	2	3	354	354-	6-	0	3	305	445-	11-	0	3	317	314-	2-	7	4	233	197-	6-	0	5	204	193-			
6-	10-	2	346	332-	6-	2	3	152	114-	6-	0	3	221	207-	11-	0	3	327	318-	2-	7	4	436	446-	3-	7	4	255	327-			
6-	11-	2	158	168-	1-	3	3	104	86-	0	7	3	410	351-	12-	0	3	206	18													

Table 3 (*cont.*)

**Table 4. Main interatomic distances and angles\***

### Anion

I-O(2) <sub>I</sub>	$= 1.992 \pm 0.012 \text{ \AA}$	O(2) <sub>I</sub> -I-O(4)	$= 88.2 \pm 0.6^\circ$
I-O(2) <sub>II'</sub>	$= 2.017 \pm 0.007$	O(4)-I-O(3)	$= 97.1 \pm 0.7$
I-O(1) <sub>I</sub>	$= 1.980 \pm 0.016$	O(3)-I-O(1) <sub>I</sub>	$= 89.6 \pm 0.7$
I-O(3)	$= 1.793 \pm 0.016$	O(1) <sub>I</sub> -I-O(2) <sub>I</sub>	$= 83.9 \pm 0.6$
I-O(4)	$= 1.810 \pm 0.015$	O(5)-I-O(4)	$= 97.6 \pm 0.6$
I-O(5) <sub>I</sub>	$= 1.818 \pm 0.010$	O(5)-I-O(3)	$= 100.0 \pm 0.7$
		O(5)-I-O(1) <sub>I</sub>	$= 87.0 \pm 0.7$
		O(5)-I-O(2) <sub>I</sub>	$= 92.0 \pm 0.6$
		O(2) <sub>II'</sub> -I-O(4)	$= 88.0 \pm 0.5$
		O(2) <sub>II'</sub> -I-O(3)	$= 89.6 \pm 0.5$
		O(2) <sub>II'</sub> -I-O(1) <sub>I</sub>	$= 86.2 \pm 0.6$
		O(2) <sub>II'</sub> -I-O(2) <sub>I</sub>	$= 77.6 \pm 0.4$

Hydrogen bonds O ··· H<sub>2</sub>O

O(5) <sub>I</sub> -O(8)	$= 2.972 \pm 0.021 \text{ \AA}$	O(5) <sub>I</sub> -O(8)-O(5)'	$= 108.9 \pm 0.8^\circ$
O(5)'-O(8)	$= 2.930 \pm 0.020$	O(4)-O(7)-O(5) <sub>III</sub>	$= 123.3 \pm 0.6$
O(4)-O(7)	$= 2.792 \pm 0.018$	O(3)-O(9)-O(5)	$= 121.1 \pm 0.7$
O(5) <sub>III</sub> -O(7)	$= 2.842 \pm 0.024$	O(4)-O(6)-O(3) <sub>III</sub>	$= 113.0 \pm 0.6$
O(3)-O(9)	$= 2.800 \pm 0.020$		
O(5)-O(9)	$= 2.854 \pm 0.026$		
O(4)-O(6)	$= 2.872 \pm 0.026$		
O(3) <sub>III</sub> -O(6)	$= 2.958 \pm 0.026$		

Coordination around K(1)

$K(1)-O(2)$	$= 2.705 \pm 0.014 \text{ \AA}$	$K(1)-O(6)_{\text{IV}}$	$= 2.860 \pm 0.016 \text{ \AA}$
$K(1)-O(3)$	$= 2.941 \pm 0.017$	$K(1)-O(4)$	$= 2.892 \pm 0.014$
$K(1)-O(3)_{\text{II}'}$	$= 2.868 \pm 0.013$	$K(1)-O(7)$	$= 3.055 \pm 0.016$
$K(1)-O(9)$	$= 3.065 \pm 0.019$	$K(1)-O(1)_{\text{III}}$	$= 2.653 \pm 0.017$

## Coordination around K(2)

$K(2)-O(8)_{IV}$	$=2.786 \pm 0.020$ Å	$K(2)-O(8)_{III'}$	$=2.749 \pm 0.020$ Å
$K(2)-O(9)$	$=2.783 \pm 0.020$	$K(2)-O(6)_{IV}$	$=3.053 \pm 0.016$
$K(2)-O(7)$	$=2.766 \pm 0.017$	$K(2)-O(6)_{III'}$	$=3.101 \pm 0.015$
$K(2)-O(9)'$	$=2.920 \pm 0.020$	$K(2)-O(7)_{III'}$	$=2.866 \pm 0.017$

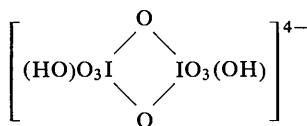
\* Asymmetric units (in fractional coordinates).

No label	$x$	$y$	$z$
'	$1-x$	$1-y$	$1-z$
I	$x-1$	$y$	$z$
II'	$1-x$	$-y$	$1-z$
III	$x$	$y$	$1+z$
III'	$1-x$	$1-y$	$2-z$
IV	$1+x$	$y$	$z$
V'	$-x$	$-y$	$1-z$

observed and calculated structure factors are compared in Table 3. The main interatomic distances and angles have been calculated (Table 4). The e.s.d.'s in the coordinates were obtained by Cruickshank's (1949) method, the e.s.d.'s of the distances by Ahmed & Cruickshank (1953) method and the e.s.d.'s of the angles by Darlow (1960) method.

### Discussion of the structure

The crystal structure of the compound (Fig. 2) is formed by anions



or written more simply  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$ , interposed between  $\text{K}^+$  cations and  $\text{H}_2\text{O}$  molecules. These water molecules, corresponding to O(6), O(7), O(8) and O(9), form a shell around the anion, to which they are bound by hydrogen bonds. Each water molecule forms two such hydrogen bonds with two anions, so giving rise to a three-dimensional network all over the structure.

The hydrogen bond distances are in the range 2.79–2.97 Å. The angles  $\text{O}-\text{H}_2\text{O}-\text{O}$  between two hydrogen bonds formed by the same water molecule have values  $109^\circ$ ,  $113^\circ$ ,  $121^\circ$  and  $123^\circ$ .

The anion  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$  has the configuration shown in Fig. 3. It can be regarded formally as obtained by condensation of two pseudo-octahedral groups  $[\text{H}_3\text{IO}_6]^{2-}$  and elimination of  $2\text{H}_2\text{O}$ , although the actual chemical process might be different; after condensation, the two octahedra have an edge in common. The two atoms O(2) and O(2'), lying on the common edge, form bridges between two I atoms. The location of the hydrogen atoms is only probable but it is supported by the following considerations. The bonds between iodine and the six neighbouring oxygen atoms are not all equivalent; three bonds, namely I-O(1), I-O(2) and I-O(2') have an average value  $\text{I}-\text{O}^*_{\text{av}} = 1.996$  Å (maximum deviation  $\pm 0.021$  Å) and the remaining three, namely I-O(3), I-O(4) and I-O(5), have an average value  $\text{I}-\text{O}_{\text{av}} = 1.807$  Å (maximum deviation  $\pm 0.011$  Å). One possible explanation of this difference is that I-O(2) and I-O(2') are longer because O(2) and O(2') are shared between two iodine atoms, while I-O(1) is longer because O(1) is actually part of a hydroxyl group, i.e. O(1) is shared between I and H.

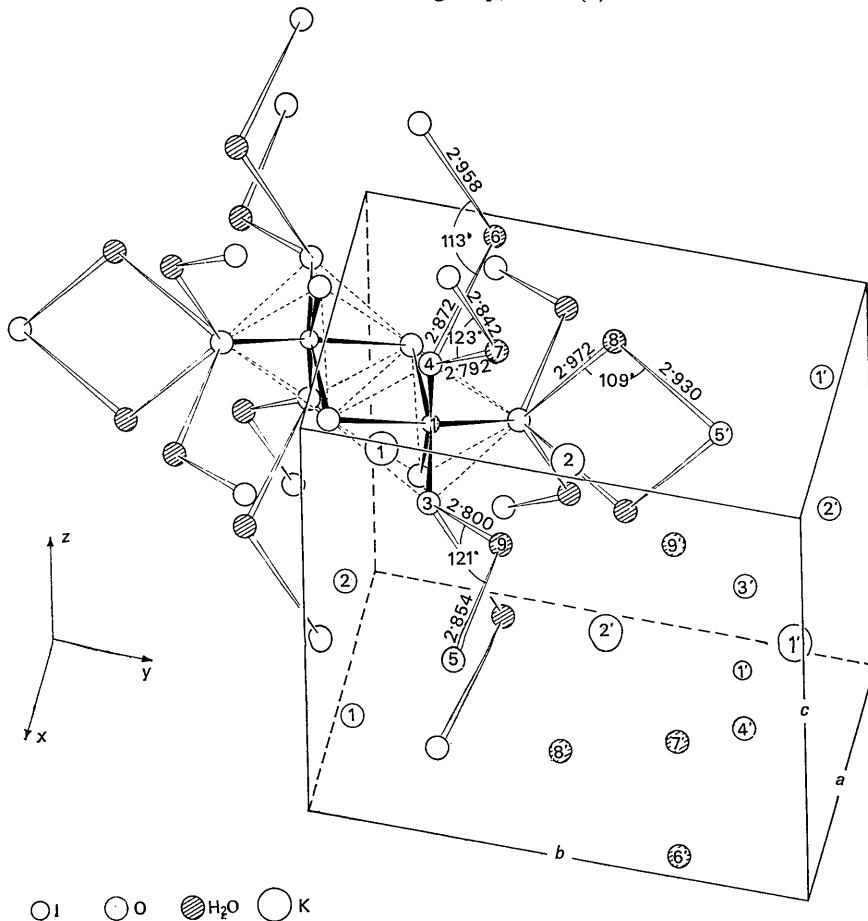


Fig. 2. Crystal structure of  $\text{K}_4\text{H}_2\text{I}_2\text{O}_{10} \cdot 8\text{H}_2\text{O}$ . Numbered atoms belong to the basic unit cell. Two octahedra forming an anion  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$  are dotted, heavy lines indicating I-O bonds. Hydrogen bond distances and angles  $\text{O}-\text{H}_2\text{O}-\text{O}$  are quoted.

The distance I–O(1), however, is longer than the corresponding distances in orthoperiodic acid,  $(\text{HO})_5\text{IO}$ , where  $\text{I}–\text{OH}_{\text{av}} = 1.89 \pm 0.02 \text{ \AA}$  (Feikema, 1963) and in  $\text{HI}_3\text{O}_8$  where  $\text{I}–\text{OH} = 1.90 \text{ \AA}$  (Feikema & Vos, 1963). The lengths of the three bonds I–O with unshared O [ $\text{I}–\text{O}(3) = 1.79 \text{ \AA}$ ,  $\text{I}–\text{O}(4) = 1.81 \text{ \AA}$  and  $\text{I}–\text{O}(5) = 1.82 \text{ \AA}$ ] are practically equal to the distance  $\text{I}–\text{O} = 1.79 \text{ \AA}$  found in the tetrahedral group  $[\text{IO}_4]^-$  (Hazlewood, 1938) and the distance  $\text{I}–\text{O} = 1.78 \text{ \AA}$  found in orthoperiodic acid (Feikema, 1963). The same distances are

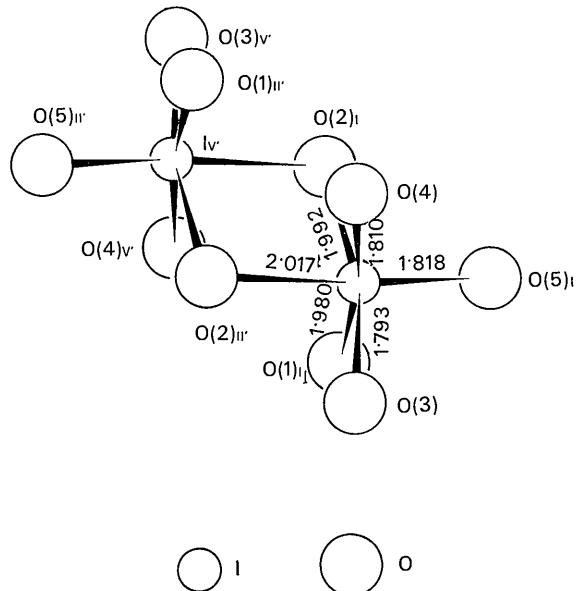


Fig. 3. The anion  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$ . The hydrogen atoms are probably bound to  $\text{O}(1)_1$  and  $\text{O}(1)_{1'}$  (for labelling of asymmetric units see Table 4).

shorter than the distances  $\text{I}–\text{O} = 1.93 \text{ \AA}$  found in the single octahedral group  $[\text{IO}_6]^{5-}$  (Helmholz, 1937). Comparison with the distances between iodine and oxygen in  $\text{HI}_3\text{O}_8$  (Feikema & Vos, 1963) shows surprising similarities in bond distances, despite the different valence of iodine. In the structure of  $\text{HI}_3\text{O}_8$ , which is an addition compound of  $\text{I}_2\text{O}_5$  and  $\text{HIO}_3$ , the distances between iodine and oxygen are  $\text{I}–\text{O}_{\text{av}} = 1.97 \text{ \AA}$  for shared oxygen in the molecule  $\text{I}_2\text{O}_5$  and  $\text{I}–\text{O}_{\text{av}} = 1.78 \text{ \AA}$  and  $\text{I}–\text{O}_{\text{av}} = 1.84 \text{ \AA}$  for unshared oxygen in  $\text{I}_2\text{O}_5$  and  $\text{HIO}_3$ , respectively.

The distances found are in agreement with a diminution of the double bond character of the bond I–O, when passing from bonds with unshared oxygen to bonds with shared oxygen atoms. The double bond character prevails in bonds I–O of the tetrahedral group  $[\text{IO}_4]^-$  and in the bond I–O of the acid  $(\text{HO})_5\text{IO}$ , but it is lowered in the bonds I–OH of the acid. The excess of double bond character causes the formation of polynuclear anions; in these the excess is eliminated at those atoms which bridge two groups and therefore the single bond character prevails in those bonds I–O where the oxygen atom is shared. The difference between bonds with shared and bonds with unshared atoms is normal, e.g. distances P–O in  $\text{Na}_4\text{P}_2\text{O}_7 \cdot 10\text{H}_2\text{O}$ , which are P–O = 1.47 Å, unshared and P–O = 1.63 Å, shared (MacArthur & Beevers, 1957). The two iodine atoms I and  $\text{I}_{1'}$  and the oxygen atoms O(3), O(5)<sub>I</sub>, O(2)<sub>I</sub>, O(5)<sub>II'</sub> and O(3)<sub>V'</sub>, all belonging to the same anion, are not exactly in the same plane. The plane of best fit (Schomaker, Waser, Marsh & Bergman, 1959) is represented by

$$1.36644x + 2.23739y + 5.53956z - 2.66978 = 0.$$

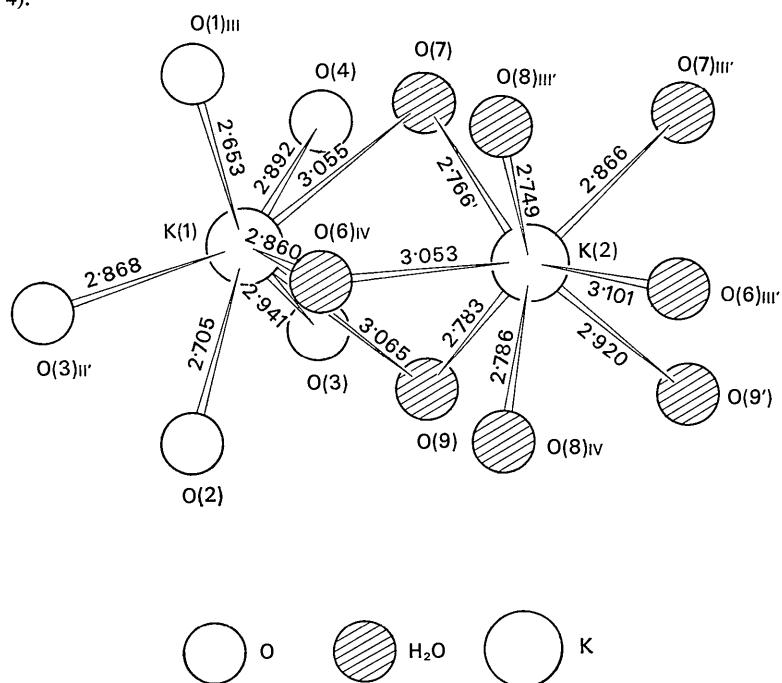


Fig. 4. Potassium ions K(1) and K(2) and surrounding atoms (for labelling of asymmetric units see Table 4).

The deviations are +0.0903 Å for I, -0.0002 Å for O(2)<sub>I</sub>, -0.0262 Å for O(3) and -0.0267 Å for O(5)<sub>I</sub> and deviations with opposite signs for centrosymmetrically related atoms of the anion. The iodine atom is significantly out of the plane and the distortion is in agreement with the fact that the bonds with respect to the plane are not equal, one bond I-OH being faced by one bond I-O.

Only the unshared oxygen atoms form hydrogen bonds with the surrounding water molecules.

The two K atoms are surrounded each by eight oxygen atoms or water molecules (Fig. 4). The coordination polyhedra differ in both form and nature for the two atoms. K(2) is surrounded by eight water molecules forming a distorted square antiprism, while K(1) is surrounded by five anionic oxygen atoms and by three water molecules. These molecules are common with the antiprism around K(2) and form a screen between the two positive ions. The distances K-O fall in the range 2.65–3.10 Å in accordance with values quoted by *International Tables for X-ray Crystallography* (1962).

The results of the structure determination show that the anion of the salt studied is correctly formulated  $[H_2I_2O_{10}]^{4-}$  and the salt should be called tetrapotassium dihydorgendecaoxodiode(VII) octahydrate (*Nomenclature of Inorganic Chemistry*, 1959). All the so-called mesoperiodates which Siebert (1960a, b) was able to identify on the basis of the characteristic infrared spectra must, whether hydrated or not, be regarded as salts of the same acid. There is a possible indication from some chromatograms prepared in this laboratory that anions derived from decaoxodiiodic (VII) acid can also exist in solution.

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