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## The Crystal Structure of Tetrapotassium Dihydrogendeca-oxodiiodate(VII) Octahydrate

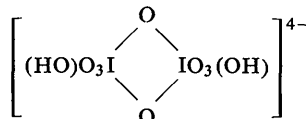
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The crystal structure of tetrapotassium dihydrogendeca-oxodiiodate (VII) octahydrate,  $K_4H_2I_2O_{10} \cdot 8H_2O$ , has been determined and refined.

The binuclear anion



can be recognized in the structure. It is formed by two octahedral groups with an edge in common, *i.e.* with two oxygen atoms shared between two iodine atoms. Distances within the binuclear anion are  $I-O_{av} = 1.996 \pm 0.021 \text{ \AA}$  for shared oxygen atoms, including that of the probable hydroxyl group, and  $I-O_{av} = 1.807 \pm 0.011 \text{ \AA}$  for unshared oxygen atoms.

The anions are interposed between  $K^+$  cations and water molecules. The water molecules form a shell around the anion to which they are bound by hydrogen bonds. Each water molecule is bound to two different anions, and a three-dimensional network of hydrogen bonds branches all over the crystal.

The potassium ions are surrounded each by eight water molecules or oxygen atoms but their environments are not equivalent.

### Introduction

The molecular and crystal structures of tetraoxoiodates(VII) (*i.e.* metaperiodates) (Hazlewood, 1938; Jaeger & Beintema, 1933) and of some orthoperiodates (Helmholz, 1937; Ferrari, Curti & Cavalca, 1944) are already known. The tetrahedral group  $[\text{IO}_4]^-$  is present in the former compounds, the octahedral group  $[\text{IO}_6]^{5-}$  in the latter. A distorted octahedral group has been found in orthoperiodic acid,  $(\text{HO})_5\text{IO}$  (Feikema, 1961, 1963). The existence of other oxoiodic(VII) acids or oxoiodates(VII) has been demonstrated or postulated by several authors (Rammelsberg, 1868; Blomstrand, 1886; Kimmins, 1889; Rae, 1931; Partington & Bahl, 1934; Pačesová & Hauptmann, 1963). In particular Siebert (1960*a, b*), from an examination of the infrared spectra of several oxoiodates(VII) infers the existence of the so-called mesoperiodic anion  $[\text{IO}_5]^{3-}$ , which should be present, for instance, in the salt  $K_2\text{HIO}_5 \cdot 4\text{H}_2\text{O}$ . The formula of this salt can also be written  $K_4\text{I}_2\text{O}_9 \cdot 9\text{H}_2\text{O}$  (Hill, 1928) or  $K_4H_2I_2O_{10} \cdot 8\text{H}_2\text{O}$ . The crystal structure of this compound has now been investigated and the last formulation has been proved to be correct.

### Experimental

Crystals of the salt  $K_4H_2I_2O_{10} \cdot 8\text{H}_2\text{O}$  can be obtained, either by evaporating or by cooling, from solutions of potassium periodate in concentrated aqueous potassium hydroxide. Its crystallographic properties were

studied by Groth (1908), who assigned them to the triclinic system, pinacoidal class.

A crystal elongated parallel to  $[001]$ , according to the orientation of Groth, has been used to determine the unit-cell constants by the roentgenographic method. The crystal has been rotated around the elongation axis, which is parallel to  $c$ . The linear constants have been measured on the photographs and the angular constants have been calculated from values given by Groth (1908) after transformation by the matrix 101/010/001. The resulting unit-cell constants are:

$$\begin{array}{ll} a = 7.06 \pm 0.01 \text{ \AA} & \alpha = 97^\circ 47' \\ b = 10.57 \pm 0.01 & \beta = 121^\circ 30' \\ c = 7.56 \pm 0.01 & \gamma = 81^\circ 55' \end{array}$$

One stoichiometric unit  $K_4H_2I_2O_{10} \cdot 8\text{H}_2\text{O}$  is contained in the unit cell. ( $\rho_{calc} = 2.500 \text{ g.cm}^{-3}$ ,  $\rho_{obs} = 2.509 \text{ g.cm}^{-3}$ ). The cell reported above can be transformed to the conventional cell (*International Tables for X-ray Crystallography*, 1959) by the matrix 101/010/100.

The space group  $P\bar{1}$  was assumed, in accordance with the morphological properties.

For the structural determinations, a crystal rotating around  $[001]$  was reduced to a cylindrical shape by filing it by a paper strip wetted by alkali solution. The intensities of the integrated reflexions  $hk0$ ,  $hk1$ ,  $hk2$ ,  $hk3$ ,  $hk4$ ,  $hk5$  have been recorded on multiple films in a Weissenberg camera and then measured by a microphotometer.  $\text{Cu } K\alpha$  radiation was used.

Lorentz and polarization factors were calculated. The transmission factors were obtained by interpolation from data given by Bond (1959) for cylindrical specimens ( $\mu R = 1.9$ ).

The atomic form factors were calculated by the Forsyth & Wells (1959) function, using the improved constants by Moore (1963). All the corrections, structure factors, three-dimensional Fourier functions, differential syntheses and anisotropic temperature factor refinement were calculated with an Olivetti Elea 6001/S computer.

#### Determination of structure

The first step of the determination of the structure was the interpretation of the Patterson projection  $P(uv)$ . Vectors I-I, K(1)-I and K(2)-I could be identified and satisfactory agreement between observed and calculated structure factors  $F(hk0)$  was obtained. The agreement was improved by calculating a few electron density projections  $\rho_0(xy)$ . The  $z$  coordinates of I, K(1) and K(2) were found approximately by generalized Patterson functions  $cP_1(uv)$  and  $sP_1(uv)$ . The remaining atoms were identified in the three-dimensional  $\rho_0(xyz)$  (Fig. 1).

The refinement started at  $R = 0.164$ . The coordinates were refined by the differential synthesis method (Booth, 1946); for the refinement of the anisotropic

thermal parameters the method described by Nardelli, Fava & Giraldi (1963) has been followed. The two refinements were applied alternately. The final atomic positions with their e.s.d.'s are reported in Table 1. (Final  $R_{hkl} = 0.099$ , observed reflexions only). The anisotropic thermal parameters are quoted in Table 2. The

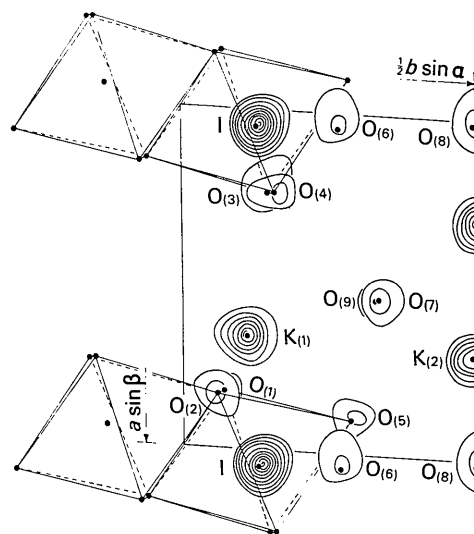


Fig. 1.  $\rho_0(xyz)$ , viewed down [001]. Composite map, arbitrary intervals. Octahedra forming binuclear anions are also drawn.

Table 1. Final atomic coordinates

Atom	$x/a$	$\sigma(x/a)$	$y/b$	$\sigma(y/b)$	$z/c$	$\sigma(z/c)$
I	0.0503	0.0001	0.1261	0.0001	0.4512	0.0002
K(1)	0.6699	0.0006	0.1117	0.0004	0.7654	0.0007
K(2)	0.6957	0.0006	0.4931	0.0004	0.8454	0.0007
O(1)	0.8300	0.0023	0.0721	0.0014	0.1601	0.0022
O(2)	0.8424	0.0014	0.0611	0.0006	0.5147	0.0018
O(3)	0.2436	0.0022	0.1445	0.0011	0.3722	0.0022
O(4)	0.2373	0.0017	0.1558	0.0010	0.7238	0.0023
O(5)	0.8976	0.0021	0.2833	0.0008	0.4025	0.0025
O(6)	0.0433	0.0023	0.2637	0.0014	0.9733	0.0029
O(7)	0.5404	0.0021	0.3341	0.0011	0.9953	0.0027
O(8)	0.0120	0.0023	0.4967	0.0011	0.7319	0.0028
O(9)	0.5414	0.0022	0.3302	0.0012	0.4844	0.0033

Table 2. Anisotropic thermal parameters\* ( $\text{\AA}^2$ )

	$B_{11}$	$B_{22}$	$B_{33}$	$B_{12}$	$B_{13}$	$B_{23}$
I	1.302	1.006	1.418	-0.112	0.488	0.137
K(1)	1.848	2.195	1.827	0.092	0.552	0.195
K(2)	2.457	1.940	2.714	-0.138	1.024	0.195
O(1)	3.092	3.236	2.035	-0.923	0.210	0.840
O(2)	1.496	1.387	2.343	0.010	0.236	0.434
O(3)	2.191	1.400	2.480	-0.759	0.703	-0.429
O(4)	2.228	1.727	3.255	0.120	0.946	0.730
O(5)	3.254	1.231	3.277	-0.931	1.484	0.186
O(6)	2.355	2.505	3.469	0.146	0.867	0.148
O(7)	2.825	1.798	3.697	-0.269	1.326	0.521
O(8)	4.977	2.582	4.869	-0.410	2.428	-0.219
O(9)	2.093	2.718	3.442	0.587	0.543	0.898

\* In the last cycle, for each atomic species I, K and O, the average shift,  $|\Delta B_{ij}|$ , and the maximum shift,  $|\Delta B_{ij}|_{\max}$ , were:

for I:  $|\Delta B_{ij}| = 0.019$  and  $|\Delta B_{ij}|_{\max} = 0.062$

for K:  $|\Delta B_{ij}| = 0.020$  and  $|\Delta B_{ij}|_{\max} = 0.060$

for O:  $|\Delta B_{ij}| = 0.031$  and  $|\Delta B_{ij}|_{\max} = 0.179$

The averages have been made over all the parameter shifts of all the atoms of the same species.

Table 3. Observed and calculated structure factors

Table with multiple columns for h, k, l, F\_o, and F\_c. The table lists observed and calculated structure factors for various Miller indices (hkl). The data is organized in several columns, with each column representing a different set of Miller indices. The observed values (F\_o) and calculated values (F\_c) are listed for each set of indices.



Table 3 (cont.)

h	k	l	10F <sub>o</sub>	10F <sub>c</sub>	h	k	l	10F <sub>o</sub>	10F <sub>c</sub>	h	k	l	10F <sub>o</sub>	10F <sub>c</sub>	h	k	l	10F <sub>o</sub>	10F <sub>c</sub>	h	k	l	10F <sub>o</sub>	10F <sub>c</sub>	h	k	l	10F <sub>o</sub>	10F <sub>c</sub>	h	k	l	10F <sub>o</sub>	10F <sub>c</sub>	
4-	2-	5	115	142	5-	3	5	322	272-	6-	4-	5	465	495-	0	6	5	358	318	1-	7-	5	235	217-	4-	8-	5	286	300	1	10-	5	245	320	
5-	2	5	---	31-	5-	3	5	223	245-	7-	4	5	311	358-	0	6	5	417	425-	2	7	5	141	144	5-	8	5	226	249	1-	10-	5	379	392	
5-	2-	5	123	21	6-	3	5	229	216-	7-	4-	5	450	464-	1	6	5	471	459	2-	7	5	331	315	5-	8	5	317	343	2-	10-	5	238	319-	
6-	2	5	310	247	6-	3-	5	535	529-	8-	4	5	165	164-	1-	6	5	315	300	2	7-	5	257	295-	6-	8	5	264	274	2-	10-	5	250	263	
6-	2-	5	---	85-	7-	3	5	160	139-	8-	4-	5	179	169-	1	6-	5	195	204-	3-	7	5	201	184-	6-	8	5	394	452	3-	10-	5	178	129-	
7-	2	5	130	101	7-	3-	5	301	306-	0	5	5	240	209	1	6-	5	855	819-	3-	7	5	277	265	7-	8	5	170	170	3-	10-	5	153	124	
7-	2-	5	195	191	8-	3	5	73	87	2	5	5	471	442	2	6	5	241	221	3-	7-	5	301	191	4-	10	5	---	---	4-	10	5	---	45	
8-	2	5	175	192	8-	3-	5	218	238-	0	5	5	402	391-	2	6	5	483	452	4-	7	5	277	262	0	9	5	217	222-	5	10	5	---	15	
8-	2-	5	280	283-	9-	3-	5	280	282-	1	5	5	438	414	2-	6	5	342	388-	4-	7	5	247	233	0	9	5	266	311	6-	10-	5	---	49-	
9-	2-	5	222	240-	1-	5	5	362	326	1	5	5	362	326	3	6	5	---	13	5	7	5	226	251	1	9	5	365	399-	6-	10-	5	---	---	
					1	5	5	206	192-	3	6	5	240	256	5	6	5	---	8	6	7	5	---	24	1	9	5	---	66-						
0	3	5	216	217-	0	4	5	149	64	1	5	5	591	549-	3	6	5	141	182-	6	7	5	314	348	1	9	5	---	23	0	11-	5	212	256	
0	3-	5	199	213	1	4	5	359	343	2	5	5	471	442	3	6	5	156	171-	7	7	5	314	356	1	9	5	447	451	1	11-	5	310	319	
1	3	5	259	250-	1	4	5	260	233-	2	5	5	257	245	4	6	5	124	101	8	7	5	221	261	2	9	5	168	172-	2	11	5	---	9-	
1	3	5	144	326-	1	4	5	300	317	2	5	5	---	86-	4	6	5	222	262-	4	6	5	---	13	2	9	5	---	104	3	11	5	232	164-	
1	3	5	350	376	1	4	5	371	303-	2	5	5	601	589-	5	6	5	---	8	5	8	5	---	37	2	9	5	267	224	4	11	5	---	62-	
1	3-	5	525	507	2	4	5	126	134	3	5	5	229	236	5	6	5	132	87	0	8	5	---	37	3	9	5	---	11-	5	11	5	270	259-	
2	3	5	143	124-	2	4	5	140	109-	3	5	5	190	167-	6	6	5	---	70-	0	8	5	104	66	3	9	5	393	350	6	11	5	323	337-	
2	3	5	461	464-	2	4	5	159	200	3	5	5	---	80	6	6	5	228	248	1	8	5	---	103-	4	9	5	267	258						
2	3-	5	520	567	2	4	5	143	120-	3	5	5	469	526-	7	6	5	102	125-	1	8	5	---	56	4	9	5	511	473						
2	3-	5	219	202	3	4	5	264	264	4	5	5	160	134-	7	6	5	94	107	1	8	5	---	65-	5	9	5	183	197	1	12	5	---	118-	
3	3	5	126	129	3	4	5	399	410-	4	5	5	215	290-	8	6	5	213	254	1	8	5	---	35	5	9	5	230	227	2	12	5	---	94	119-
3	3	5	563	626-	3	4	5	211	224	5	5	5	124	104-	2	8	5	139	183-	2	8	5	---	65	6	9	5	197	193	3	12	5	282	322-	
3	3-	5	346	374	3	4	5	410	486-	5	5	5	185	192-	2	8	5	---	65	7	9	5	---	65	7	9	5	138	131	4	12	5	298	332-	
3	3-	5	245	287-	4	4	5	757	730-	6	5	5	315	334-	0	7	5	203	200	2	8	5	169	194-	0	10	5	---	230-	5	12	5	---	242-	
4	3	5	---	140	4	4	5	575	646-	6	5	5	278	285-	0	7	5	113	134	1	8	5	---	103-	4	9	5	267	258						
4	3	5	465	498-	5	4	5	388	354-	7	5	5	194	231-	1	7	5	177	158	3	8	5	---	252	3	8	5	252	285						
4	3-	5	169	247	5	4	5	322	364-	7	5	5	126	101-	1	7	5	499	490	3	8	5	524	469	0	10	5	323	355						
4	3-	5	---	39-	6	4	5	243	264-	8	5	5	108	122	1	7	5	303	334-	4	8	5	392	347	1	10	5	310	341-						

Table 4. Main interatomic distances and angles\*

Anion			
I-O(2) <sub>I</sub>	= 1.992 ± 0.012 Å	O(2) <sub>I</sub> -I-O(4)	= 88.2 ± 0.6°
I-O(2) <sub>II'</sub>	= 2.017 ± 0.007	O(4)-I-O(3)	= 97.1 ± 0.7
I-O(1) <sub>I</sub>	= 1.980 ± 0.016	O(3)-I-O(1) <sub>I</sub>	= 89.6 ± 0.7
I-O(3)	= 1.793 ± 0.016	O(1) <sub>I</sub> -I-O(2) <sub>I</sub>	= 83.9 ± 0.6
I-O(4)	= 1.810 ± 0.015	O(5) <sub>I</sub> -I-O(4)	= 97.6 ± 0.6
I-O(5) <sub>I</sub>	= 1.818 ± 0.010	O(5) <sub>I</sub> -I-O(3)	= 100.0 ± 0.7
		O(5) <sub>I</sub> -I-O(1) <sub>I</sub>	= 87.0 ± 0.7
		O(5) <sub>I</sub> -I-O(2) <sub>I</sub>	= 92.0 ± 0.6
		O(2) <sub>II'</sub> -I-O(4)	= 88.0 ± 0.5
		O(2) <sub>II'</sub> -I-O(3)	= 89.6 ± 0.5
		O(2) <sub>II'</sub> -I-O(1) <sub>I</sub>	= 86.2 ± 0.6
		O(2) <sub>II'</sub> -I-O(2) <sub>I</sub>	= 77.6 ± 0.4
Hydrogen bonds O...H <sub>2</sub> O			
O(5) <sub>I</sub> -O(8)	= 2.972 ± 0.021 Å	O(5) <sub>I</sub> -O(8)-O(5)'	= 108.9 ± 0.8°
O(5)'-O(8)	= 2.930 ± 0.020	O(4)-O(7)-O(5) <sub>III</sub>	= 123.3 ± 0.6
O(4)-O(7)	= 2.792 ± 0.018	O(3)-O(9)-O(5)	= 121.1 ± 0.7
O(5) <sub>III</sub> -O(7)	= 2.842 ± 0.024	O(4)-O(6)-O(3) <sub>III</sub>	= 113.0 ± 0.6
O(3)-O(9)	= 2.800 ± 0.020		
O(5)-O(9)	= 2.854 ± 0.026		
O(4)-O(6)	= 2.872 ± 0.026		
O(3) <sub>III</sub> -O(6)	= 2.958 ± 0.026		
Coordination around K(1)			
K(1)-O(2)	= 2.705 ± 0.014 Å	K(1)-O(6) <sub>IV</sub>	= 2.860 ± 0.016 Å
K(1)-O(3)	= 2.941 ± 0.017	K(1)-O(4)	= 2.892 ± 0.014
K(1)-O(3) <sub>II'</sub>	= 2.868 ± 0.013	K(1)-O(7)	= 3.055 ± 0.016
K(1)-O(9)	= 3.065 ± 0.019	K(1)-O(1) <sub>III</sub>	= 2.653 ± 0.017
Coordination around K(2)			
K(2)-O(8) <sub>IV</sub>	= 2.786 ± 0.020 Å	K(2)-O(8) <sub>III'</sub>	= 2.749 ± 0.020 Å
K(2)-O(9)	= 2.783 ± 0.020	K(2)-O(6) <sub>IV</sub>	= 3.053 ± 0.016
K(2)-O(7)	= 2.766 ± 0.017	K(2)-O(6) <sub>III'</sub>	= 3.101 ± 0.015
K(2)-O(9)'	= 2.920 ± 0.020	K(2)-O(7) <sub>III'</sub>	= 2.866 ± 0.017

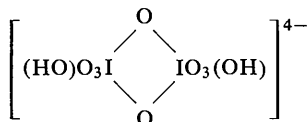
\* Asymmetric units (in fractional coordinates).

No label	x	y	z
I	1-x	1-y	1-z
II'	x-1	-y	z
III	1-x	-y	1-z
III'	x	y	1+z
IV	1-x	1-y	2-z
V'	1+x	y	z
	-x	-y	1-z

observed and calculated structure factors are compared in Table 3. The main interatomic distances and angles have been calculated (Table 4). The e.s.d.'s in the coordinates were obtained by Cruickshank's (1949) method, the e.s.d.'s of the distances by Ahmed & Cruickshank (1953) method and the e.s.d.'s of the angles by Darlow (1960) method.

### Discussion of the structure

The crystal structure of the compound (Fig. 2) is formed by anions



or written more simply  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$ , interposed between  $\text{K}^+$  cations and  $\text{H}_2\text{O}$  molecules. These water molecules, corresponding to O(6), O(7), O(8) and O(9), form a shell around the anion, to which they are bound by hydrogen bonds. Each water molecule forms two such hydrogen bonds with two anions, so giving rise to a three-dimensional network all over the structure.

The hydrogen bond distances are in the range 2.79–2.97 Å. The angles  $\text{O}-\text{H}_2\text{O}-\text{O}$  between two hydrogen bonds formed by the same water molecule have values  $109^\circ$ ,  $113^\circ$ ,  $121^\circ$  and  $123^\circ$ .

The anion  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$  has the configuration shown in Fig. 3. It can be regarded formally as obtained by condensation of two pseudo-octahedral groups  $[\text{H}_3\text{IO}_6]^{2-}$  and elimination of  $2\text{H}_2\text{O}$ , although the actual chemical process might be different; after condensation, the two octahedra have an edge in common. The two atoms O(2) and O(2'), lying on the common edge, form bridges between two I atoms. The location of the hydrogen atoms is only probable but it is supported by the following considerations. The bonds between iodine and the six neighbouring oxygen atoms are not all equivalent; three bonds, namely I–O(1), I–O(2) and I–O(2') have an average value  $\text{I}-\text{O}^*_{\text{av}} = 1.996 \text{ \AA}$  (maximum deviation  $\pm 0.021 \text{ \AA}$ ) and the remaining three, namely I–O(3), I–O(4) and I–O(5), have an average value  $\text{I}-\text{O}_{\text{av}} = 1.807 \text{ \AA}$  (maximum deviation  $\pm 0.011 \text{ \AA}$ ). One possible explanation of this difference is that I–O(2) and I–O(2') are longer because O(2) and O(2') are shared between two iodine atoms, while I–O(1) is longer because O(1) is actually part of a hydroxyl group, *i.e.* O(1) is shared between I and H.

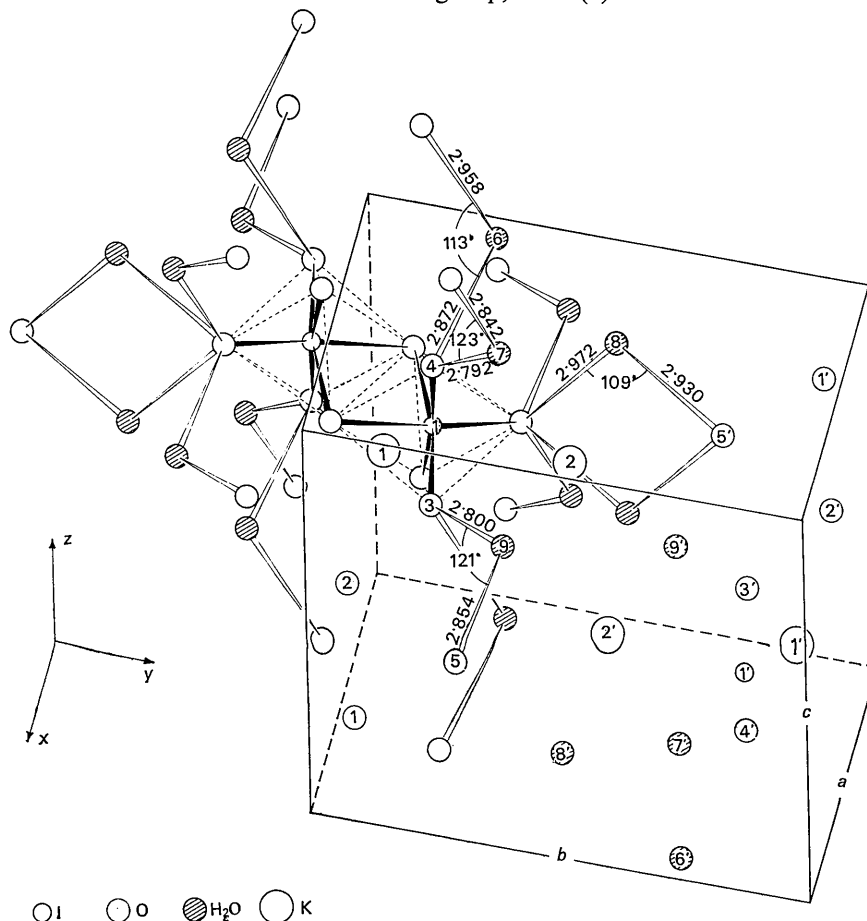


Fig. 2. Crystal structure of  $\text{K}_4\text{H}_2\text{I}_2\text{O}_{10} \cdot 8\text{H}_2\text{O}$ . Numbered atoms belong to the basic unit cell. Two octahedra forming an anion  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$  are dotted, heavy lines indicating I–O bonds. Hydrogen bond distances and angles  $\text{O}-\text{H}_2\text{O}-\text{O}$  are quoted.

The distance I-O(1), however, is longer than the corresponding distances in orthoperiodic acid,  $(\text{HO})_5\text{IO}$ , where  $\text{I-OH}_{\text{av}} = 1.89 \pm 0.02 \text{ \AA}$  (Feikema, 1963) and in  $\text{HI}_3\text{O}_8$  where  $\text{I-OH} = 1.90 \text{ \AA}$  (Feikema & Vos, 1963). The lengths of the three bonds I-O with unshared O [ $\text{I-O}(3) = 1.79 \text{ \AA}$ ,  $\text{I-O}(4) = 1.81 \text{ \AA}$  and  $\text{I-O}(5) = 1.82 \text{ \AA}$ ] are practically equal to the distance  $\text{I-O} = 1.79 \text{ \AA}$  found in the tetrahedral group  $[\text{IO}_4]^-$  (Hazlewood, 1938) and the distance  $\text{I-O} = 1.78 \text{ \AA}$  found in orthoperiodic acid (Feikema, 1963). The same distances are

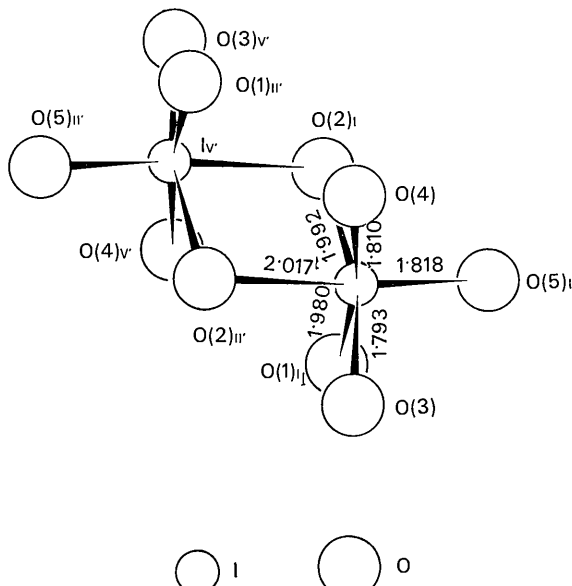


Fig. 3. The anion  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$ . The hydrogen atoms are probably bound to  $\text{O}(1)_{\text{II}}$  and  $\text{O}(1)_{\text{II}'}$  (for labelling of asymmetric units see Table 4).

shorter than the distances  $\text{I-O} = 1.93 \text{ \AA}$  found in the single octahedral group  $[\text{IO}_6]^{5-}$  (Helmholz, 1937). Comparison with the distances between iodine and oxygen in  $\text{HI}_3\text{O}_8$  (Feikema & Vos, 1963) shows surprising similarities in bond distances, despite the different valence of iodine. In the structure of  $\text{HI}_3\text{O}_8$ , which is an addition compound of  $\text{I}_2\text{O}_5$  and  $\text{HIO}_3$ , the distances between iodine and oxygen are  $\text{I-O}_{\text{av}} = 1.97 \text{ \AA}$  for shared oxygen in the molecule  $\text{I}_2\text{O}_5$  and  $\text{I-O}_{\text{av}} = 1.78 \text{ \AA}$  and  $\text{I-O}_{\text{av}} = 1.84 \text{ \AA}$  for unshared oxygen in  $\text{I}_2\text{O}_5$  and  $\text{HIO}_3$ , respectively.

The distances found are in agreement with a diminution of the double bond character of the bond I-O, when passing from bonds with unshared oxygen to bonds with shared oxygen atoms. The double bond character prevails in bonds I-O of the tetrahedral group  $[\text{IO}_4]^-$  and in the bond I-O of the acid  $(\text{HO})_5\text{IO}$ , but it is lowered in the bonds I-OH of the acid. The excess of double bond character causes the formation of polynuclear anions; in these the excess is eliminated at those atoms which bridge two groups and therefore the single bond character prevails in those bonds I-O where the oxygen atom is shared. The difference between bonds with shared and bonds with unshared atoms is normal, e.g. distances P-O in  $\text{Na}_4\text{P}_2\text{O}_7 \cdot 10\text{H}_2\text{O}$ , which are  $\text{P-O} = 1.47 \text{ \AA}$ , unshared and  $\text{P-O} = 1.63 \text{ \AA}$ , shared (MacArthur & Beevers, 1957). The two iodine atoms I and  $\text{I}'$  and the oxygen atoms  $\text{O}(3)$ ,  $\text{O}(5)_{\text{I}}$ ,  $\text{O}(2)_{\text{I}}$ ,  $\text{O}(2)_{\text{II}'}$ ,  $\text{O}(5)_{\text{II}'}$  and  $\text{O}(3)_{\text{V}'}$ , all belonging to the same anion, are not exactly in the same plane. The plane of best fit (Schomaker, Waser, Marsh & Bergman, 1959) is represented by

$$1.36644x + 2.23739y + 5.53956z - 2.66978 = 0.$$

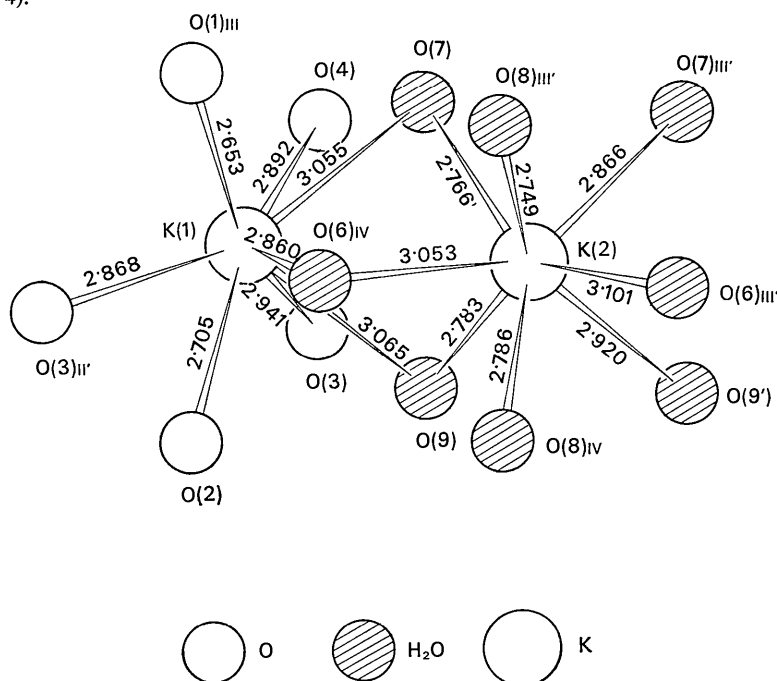


Fig. 4. Potassium ions  $\text{K}(1)$  and  $\text{K}(2)$  and surrounding atoms (for labelling of asymmetric units see Table 4).

The deviations are  $+0.0903 \text{ \AA}$  for I,  $-0.0002 \text{ \AA}$  for O(2)<sub>I</sub>,  $-0.0262 \text{ \AA}$  for O(3) and  $-0.0267 \text{ \AA}$  for O(5)<sub>I</sub> and deviations with opposite signs for centrosymmetrically related atoms of the anion. The iodine atom is significantly out of the plane and the distortion is in agreement with the fact that the bonds with respect to the plane are not equal, one bond I-OH being faced by one bond I-O.

Only the unshared oxygen atoms form hydrogen bonds with the surrounding water molecules.

The two K atoms are surrounded each by eight oxygen atoms or water molecules (Fig. 4). The coordination polyhedra differ in both form and nature for the two atoms. K(2) is surrounded by eight water molecules forming a distorted square antiprism, while K(1) is surrounded by five anionic oxygen atoms and by three water molecules. These molecules are common with the antiprism around K(2) and form a screen between the two positive ions. The distances K-O fall in the range  $2.65\text{--}3.10 \text{ \AA}$  in accordance with values quoted by *International Tables for X-ray Crystallography* (1962).

The results of the structure determination show that the anion of the salt studied is correctly formulated  $[\text{H}_2\text{I}_2\text{O}_{10}]^{4-}$  and the salt should be called tetrapotassium dihydrogendecaoxodiiiodate(VII) octahydrate (*Nomenclature of Inorganic Chemistry*, 1959). All the so-called mesoperiodates which Siebert (1960*a, b*) was able to identify on the basis of the characteristic infrared spectra must, whether hydrated or not, be regarded as salts of the same acid. There is a possible indication from some chromatograms prepared in this laboratory that anions derived from decaoxodiiiodic (VII) acid can also exist in solution.

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